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Ratios Questions 1. Aluminum reacts with
oxygen to produce aluminum oxide as follows:
 $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ a. If you use 2.3 moles
of Al, how many moles of Al_2O_3 can you
make? b. If you want 3.9 moles of Al_2O_3 ,
how many moles of O_2 are needed? 2.

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Chapter 12: Stoichiometry Mole Ratios

Questions 1. Aluminum reacts with oxygen to produce aluminum oxide as follows: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
a. If you use 2.3 moles of Al, how many moles of Al_2O_3 can you make? b. If you want 3.9 moles of Al_2O_3
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points TTheStoichiometry section provides
answers to the following challenges: The
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answers to equation worksheet answers for the research section of the Biology Math Guide.

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Stoichiometry127 SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles,

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