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~~Equilibrium Answer~~
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Answers Chapter 5 Concept

Review: Measurements and

Calculations in Chemistry 1.

Accuracy is the extent to which a measurement approaches the true value of a quantity; precision is the extent to which a series

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Answer Key

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At 450°C ,
the value of the equilibrium
constant for the following system
is 6.59×10^{-3} . If $[\text{NH}_3] = 1.23$

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Answer Key

$1 \times 10^{-4} \text{ M}$ and $[\text{H}_2] = 2.75 \times 10^{-2} \text{ M}$ at equilibrium, determine the concentration of N_2 at that point.

$$\text{N}_2(\text{g}) + 3 \text{H}_2(\text{g})$$

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Answer Key 2



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The state of equilibrium can be maintained over time as long as all factors remain the same. The state

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~~Answer Key~~ of equilibrium can be maintained over time as long as all factors change. The chemical...

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Equilibrium Answer At 450 C,
the value of the equilibrium
constant for the following system
is 6.59×10^{-3} . If $[\text{NH}_3] = 1.23$
 $\times 10^{-4}\text{M}$ and $[\text{H}_2] = 2.75 \times 10^{-2}\text{M}$
at equilibrium, determine
the concentration of N_2 at that

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point. $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$

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a reversible reaction in which
there is no longer any change in

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the concentration of reactants and products. chemical equilibrium. a state of balance in which the rate of a forward reaction equals the rate of the reverse reaction and the concentrations of products and reactants remain unchanged. equilibrium.

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Answer Key

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Answers REVIEW Chemical

Equilibrium SECTION 4 SHORT

ANSWER Answer the following

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Answers in the space provided. 1.
Match the solution type on the right to the corresponding relationship between the ion product and the K_{sp} for that solution, listed on the left. _____
The ion product ...

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~~Chapter 18 Review Chemical Equilibrium Answers~~

Use the following chemical equation to answer the question below:

$$\text{Mg}(s) + 2\text{H}^+ (aq) + 2\text{Cl}^- (aq) \rightarrow \text{Mg}^{2+} (aq) + 2\text{Cl}^- (aq) + \text{H}_2 (g)$$

$\text{H}_2\text{O}(l)$ If 0.048 g of magnesium completely reacts in 20 s, what is

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the average reaction rate in
moles/second over that time
interval? Average rate 9.9×10^{-5}
mol/s MODERN CHEMISTRY
REACTION KINETICS 141
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~~R k H 2 NO 2 3 Use the following
chemical equation to ...~~

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ANS: K_p (equilibrium constant) is

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Answer Key independent of pressure and concentration. 34 .One mole of a compound AB reacts with one mole of a compound CD according to the equation $AB + CD \rightleftharpoons AD + CB$. When equilibrium had been established it was found that 34 mole each of reactant AB and CD

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Answer Key
had been converted to AD and CB.

~~Chemical Equilibrium Important
Questions And Answers~~

The Chemical Equilibrium chapter of this Holt McDougal Modern Chemistry Companion Course helps students learn the essential

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answers chapter 4. stances. 2.
Answers will vary but may include

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iron rusting, milk. . .

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Answers Chapter 4~~

18) The following equilibrium is readily established: $\text{SO}_2\text{Cl}_2 (\text{g}) \rightleftharpoons \text{SO}_2 (\text{g}) + \text{Cl}_2 (\text{g})$ At equilibrium at 373 K, a 1.00-L reaction vessel

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contains 0.0106 mol of SO_2Cl_2 and 0.0287 mol each of SO_2 and Cl_2 .

What is K_{eq} for the reaction at 373 K? A) 12.8 B) 2.72 C) 0.0781 D) 2.39 E) 0.418

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