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Answer the following questions in the space provided. 1. \_\_\_\_\_ The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L. 2.

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Three of the primary components of air are carbon dioxide, nitrogen, and oxygen. In a sample containing a mixture of only these gases at exactly 1 atm, the partial pressures of carbon dioxide and nitrogen are given as  $P_{\text{CO}_2} = 0.285 \text{ torr}$  and  $P_{\text{N}_2} = 593.525 \text{ torr}$ . What is the partial pressure of oxygen?

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SECTION 1 continued 5. Convert a  
pressure of 0.200 atm to the following  
units: 152 a. mm Hg 20.3 b. kPa 6.

When an explosive like TNT is  
detonated, a mixture of gases at high  
temperature is created. Suppose that  
gas X has a pressure of 50 atm, gas Y  
has a pressure of 20 atm, and gas Z  
has a pressure of 10 atm. 80 atm a.  
What is the total ...

*mc06se cFMs r i-vi*

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CHAPTER 11 REVIEW Gases

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## SECTION 4 SHORT ANSWER

Answer the following questions in the space provided. 1. List the following gases in order of increasing rate of effusion.

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CHAPTER 11 REVIEW Gases

## SECTION 3 SHORT ANSWER

Answer the following questions in the space provided. 1. \_\_\_\_\_ The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (b) multiplied by 22.4 L. (c) divided by the mass of 1 mol. (d) divided ...

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letter of the term or phrase that best  
completes each sentence or best  
answers each question. \_\_\_\_\_ 1. At  
the same temperature and pressure,  
balloons of equal volume always  
contain

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